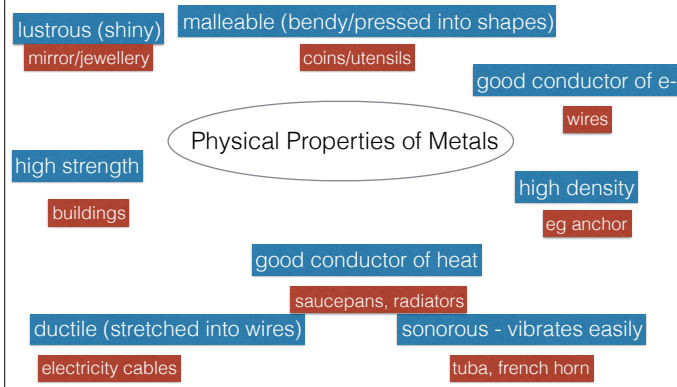
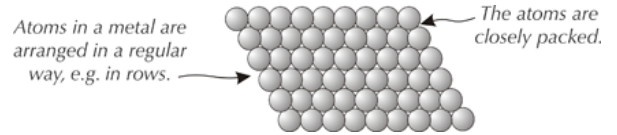


Properties of metals



Structure of Metals

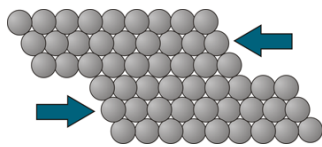
Why are metals malleable?



1. The atoms in a metal are arranged in a regular pattern.
2. The atoms are in layers.
3. Layers can slide over each other.

Structure of Metals

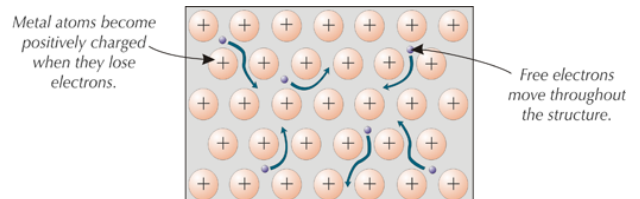
Why are metals malleable?



1. The atoms in a metal are arranged in a regular pattern.
2. The atoms are in layers.
3. Layers can slide over each other.

Structure of Metals

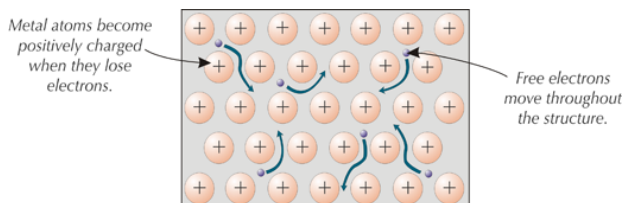
Why do metals have a high boiling point?



1. In metals, the electrons in the outer shells of the atoms are delocalised.
2. There are strong forces of electrostatic attraction between the positive metal ions and the negative electrons.
3. These strong forces hold the metal structure together.

Structure of Metals

Why do metals have a high electrical conductivity?



1. In metals, the electrons in the outer shells of the atoms are delocalised.
2. They're free to move through the whole structure
3. They carry the electric current (or heat energy).

Practice Questions — Fact Recall

1. Describe the structure of a metal.
[Show answer](#)
2. **H** What are delocalised electrons?
[Show answer](#)
3. a. **H** What type of forces hold the particles in a metal together?
[Show answer](#)
b. **H** Why do these forces exist in a metal?
[Show answer](#)