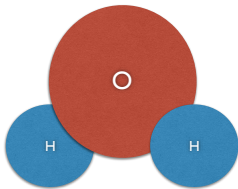


Molecular formula



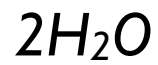
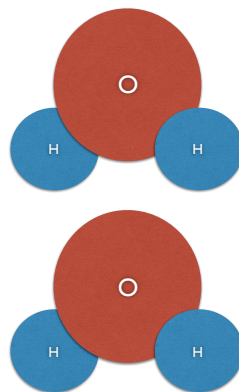
The molecular formula tells us how many atoms are in a compound.

In this molecule are 2 atoms of hydrogen (H) and one atom of oxygen (O)

We write this as H_2O

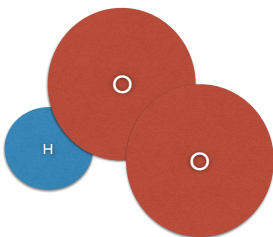
this number has to be small

Common mistakes



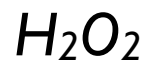
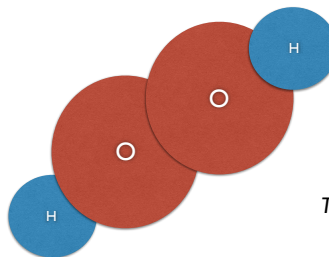
a number here means 2 molecules of water

Common mistakes



a number here means 2 atoms of oxygen - a different compound

Common mistakes





This would also be a different chemical, because it would have two oxygen atoms and two hydrogen atoms.

Molecular formula



KEY

 hydrogen atoms

 oxygen atoms


 carbon atoms

 nitrogen atoms

Can you give the Molecular formula?

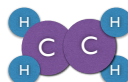
KEY

 hydrogen atoms

 oxygen atoms

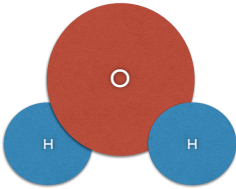
 carbon atoms

 nitrogen atoms



Naming compounds

Step 1 - identify the atoms



Naming compounds

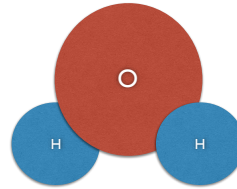
Step 1 - identify the atoms

H = hydrogen

a compound with hydrogen is hydrous or a hydride

O = oxygen

a compound with oxygen is an oxide



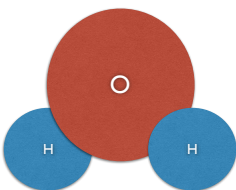
Naming compounds

Step 2 - count the atoms

1 oxygen atom

2 hydrogen atoms

2 atoms is called di-



Naming compounds

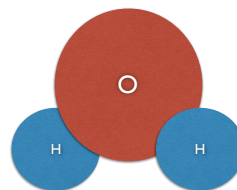
Step 3 - put it all together

di hydrous oxide

means 2 hydrogen atoms

& 1 oxygen atom

the common name for this compound is water



NAME THE COMPOUND

KEY



hydrogen gas
(element)



hydrogen atoms



water
(di-hydrous oxide)



oxygen atoms



carbon dioxide



carbon atoms



ammonia
(tri-hydrous nitride)



nitrogen atoms



methane

Remember

mon - 1

di - 2

tri - 3